Spectrophotometric determination of selenium in various samples using 5,6-diaminouracil hydrochloride (DAUH)

Salwa Fares Raasi

Department of Chemistry, Faculty of Sciences, University of Al- Baath Homs
Mobile: 00963-966-243153

E-mail address: salwararassi@gmail.com

Keywords: Selenium determination, spectrophotometry, 5,6-diaminouracil hydrochloride (DAUH).

ABSTRACT: A simple, rapid and sensitive spectrophotometric method was developed for the determination of Se (IV) using 5,6-diaminouracil hydrochloride (DAUH) as an analytical reagent. The reagent has been synthesized and characterized using IR, 1H NMR. The metal ion in aqueous medium forms yellow colored complex with DAUH showing maximum absorbance at 343 nm. Hence, analytical studies were further carried out at 343 nm. The reagent reacts with selenium in acidic medium to form yellow colored 1:1 (M:L) complex. The color reactions are instantaneous and absorbance values remain constant for one week. The composition of the Se (IV) complex with DAUH was studied by the method of job’s continuous variation and molar ratio method. Beer’s law was obeyed in the range 0.2-15.0 μg mL⁻¹ of Se (IV). The molar absorptivity and Sandell’s sensitivity of the method were found to be 0.5×10⁴ L.mol⁻¹.cm⁻¹ and 0.031μg.cm² respectively. Since DAUH method is more sensitive, it was applied for the determination of selenium in samples.

1. INTRODUCTION

Spectrophotometric methods is one of the most important ways, which is used in analytical chemistry, because it wide used in researches field, and has sensitivity, accuracy, simplify and low cost. The diversity UV-Vis. analysis (direct, extraction, kinetic…etc) by using organic reagents make it possible to determine a lot of elements in many analytical samples

Selenium is an essential trace nutrient and selenium deficiency diseases are well known in veterinary medicine. Above trace levels, ingested selenium is toxic to animals and may be toxic to humans. The selenium concentration of most drinking waters and natural waters is less than 10 μg mL⁻¹. Selenium is widely distributed in nature, in relatively small concentrations in rocks, plants, coal and other fossil fuels. Selenium enters into natural water through seepage from splendiferous soil and industrial waste. Selenium compounds have extensive applications and because of its significance, several analytical techniques have been reported for the determination of selenium, which includes spectro-fluorometry, electro thermal atomic absorption spectrometry, hydride generation, polarography, cathode-stripping voltammetry, radiochemical neutron activation analysis, and flow- injection techniques [1]. There are many reagents available for the spectrophotometric determination of selenium, among which, the recently used were J-acid [2], Leuco crystal violet [3], resazurin [4], sodium salt of hexamethyleneimine carbodithioate [5], 1-naphthylamine-7-sulphonic acid [6], variamine blue [7]. When selenium is present in animal feeds at a concentration less than 0.1 mg L⁻¹, deficiency symptoms develop, but when present at a higher concentration, exceeding 5 mgL⁻¹, chronic selenosis occurs. Selenium tends to weaken the toxic action of some heavy metals in animal and human organisms [8-9].

Selenium is widely distributed in the environment (waters, soil, and air) albeit generally in very low concentrations (≤1μg/g). Selenium is a naturally occurring element found in rocks[10], soil, water, air and animals. Selenium is a trace mineral that is essential to good health but required only in small amounts [11-13]. Selenium is incorporated into proteins to make selenoproteins, which are important antioxidant enzymes. The antioxidant properties of selenoproteins help to prevent cellular damage from free radicals. Other selenoproteins help regulate thyroid function and play a role in the immune system [14-16]. Selenium compound are widely used in paints, dyes, glass electrical, rubber, insecticides, and many other industries. Some industrial and agricultural
processes release selenium as a by-product and selenium from such sources has caused environmental disaster [17]. The threshold limit value for selenium compounds in air is 0.1-0.2 mg dm³ in water it is 4.0 ppm. The toxicity, availability and environmental mobility of selenium are very much dependent on its chemical form [18]. Selenium can occur in different oxidation states in organic and inorganic compounds. In many environmental matrices, e.g. natural water and soils, the predominant oxidation state of selenium are Se (IV) and Se (VI). Selenium is also reported to be present in cigarette paper, tobacco [19] and various cosmetic samples [20]. Because of its significance, several analytical techniques have been reported concerning the determination of selenium [21-24].

Plant foods are the major dietary sources of selenium in most countries throughout the world. The content of selenium in food depends on the selenium content of the soil where plants are grown or animals are raised.

In the present study, we are reporting rapid, simple, sensitive and selective methods for the determination of traces of selenium (IV) with DAUH, anew reagent. This paper describes synthesis, characterization and analytical properties of new reagent (DAUH). Since the reagent is more sensitive, it is used for the determination of selenium in various samples.

2. EXPERIMENTAL
2.1. Apparatus
A Jasco V–530 UV–VIS spectrophotometer (Japan) with 1 cm quartz cells was used for all absorbance measurements under the following operating conditions: scan speed medium (400 nm/min), scan range 200–1100 nm and slit width 2 nm. Spectra were automatically obtained by Jasco system software. pH measurements were made with ORION 250A (USA) with combined glass pH electrode.

NMR Spectrometry Bruker 400MHz, FT-IR 4100 (Fourier transform infrared spectrometer) Jasco the results of the suggested method were coincidental with the analysis data of the same samples with the Hydride Generation-Atomic Absorption technique as a comparative method

2.2. Reagents and materials
All chemicals used were of analytical-reagent grade of the highest purity available procured from Merck. Doubly distilled de-ionized water was used throughout the experiment.

2.3. Preparation of reagent (DAUH)
DAUH was prepared In flask equipped with a reflux condenser and an efficient stirrer is placed 500 mL of absolute (99.8%) ethanol. To this is added 19.7 g. (1.72 g. atom) of sodium, and, after solution is complete, 45.57 ml.(48.65 g., 0.86 mole) of ethyl cyanoacetate and 50.60 g. (0.86 mole) of urea are added. The mixture is heated 4 hours, At the end of the reaction time, 500 mL of hot (80°) water is added., the stirred mixture is heated at 80° for 15 minutes and is then neutralized. Additional glacial acetic acid (37.5 ml.) is added(STEP(1), followed by cautious addition of a solution of 32.4 g. (0.94 mole) of sodium nitrite dissolved in 35 ml. of water. The dark red nitroso compound separates almost immediately.

The slurry is stirred while being heated on a steam bath, and solid sodium hydrosulfite is added until the red color of the nitroso compound is completely bleached. Then an additional 15g of sodium hydrosulfite is added; the light tan suspension is stirred with heating for 15 minutes more and is allowed to cool. The dense 5,6-diamino- Uracil bisulfite is filtered from the cooled solution, washed well with water, and partially dried.

Concentrated hydrochloric acid is added until the consistency of the resulting mixture is such as to permit mechanical stirring (50 to 100 ml. of acid). The slurry is heated on a steam bath with stirring for 1 hour. The tan 5,6-diamino Uracil hydrochloride is filtered, washed well with acetone, and vacuum-dried over phosphorus pentoxide. The yield of 5,6-diamino- Uracil hydrochloride is. (90 %). (Figure 1).
2.4. Characterization of reagent DAUH

The reagent has been synthesized and characterized by IR, 1HNMR data. Infrared spectrum of DAUH shows bands at 3409 , 3404 - 3285, 1695, 1715, 1665, 1197, 1329, respectively corresponding to $\nu$ (O-H) $\nu$ (NH$_2$), (C= N), $\nu$ (C= O), (C = C) ( C - C), (C - N )

H1NMR spectrum of DAUH (CDCl$_3$+DMSO) showed signals at 8.00 (6H), 11.53 (1H), 7.83(1H), due to C5-NH$_3^+$, C6-NH$_3^+$, C1-OH, C3-OH. The molecular formula of the reagent is C$_4$H$_8$N$_4$O$_2$Cl$_2$ (M.Wt, 187).(Figure 2).

2.5. pKa values of reagents

The pKa values were determined by recording the UV-Visible spectra of $1 \times 10^{-4}$M solutions of the reagent at various pH values and by taking the arithmetic mean of the values obtained from the measurements at different wave lengths determined spectrophotometrically depended on reference analytical methods (half height, Limit absorbance and Colleter's). The values of deprotonation of DAUH were (4.63, 8.14 and 10.62)

2.6. DAUH solution

A $1 \times 10^{-2}$M solution was prepared by dissolving 0.0187 g of DAUH in 100 ml of water containing few drops of conc.HCl, The reagent solution is stable for at least 24 h.

2.7. Se (IV) solution

A $1 \times 10^{-2}$M stock solution of selenium was prepared by dissolving 0.1890 g of sodium selenate Na$_2$Seo4 (Merck Darmstadt) in double distilled water containing few drops of conc.H$_2$SO$_4$ and made up to the mark in a100 mL volumetric flask. Aliquots of this solution were standardized with EDTA titration using xylenol orange as an indicator. Dilute solutions were prepared from this stock solution.

2.8. Procedure for Preparation of water samples

Different water samples were collected from various places. The samples (150 ml) were stored at 5oC in metal free polyethylene bottles. Water samples were filtered through whatman filter paper no. 41 and collected into 250 ml beakers. All the filtered environmental water samples were evaporated nearly to dryness with a mixture of 10 ml con HNO$_3$ and 5ml of con H$_2$SO$_4$ and then cooled to room temperature. The sample was digested in the presence of an excess potassium permanganate solution according to the method recommended by Fifiled et al.,5. The residues were
then heated with 10 ml of deionized water in order to dissolve the salts. The solutions were cooled
and neutralized with dilute NH₄OH. The digest was transferred into a 25 ml calibrated flask and
diluted up to the mark with deionized water.

2.9. Procedure for Preparation of various samples

To determine selenium in Yellow Corn, Peas, Chick-pea, Wheat, Almond. Takes 1-4 grams of
the sample after crushed and put in Arlanmayr hydride generating device, and add to 20 ml of
distilled water, then slowly mix, were heated on electric heater for 10 minutes while stirring, then
cooled resulting solution and add 10 mL of tin chloride solution SnCl₂.2H₂O (0.33% HCl in the
center) and 5 ml of potassium iodide solution (0.15%), leaves the mixture for 5 minutes, add 10 ml
of alcohol (to reduce the foam produced during the reaction), then add 8 grams metal of zinc, device
Closing to collect hydride in another Arlanmayr that contains 2 mL 0.5M NaOH and 2% water
Oxygen, leaves the reaction for an hour to two hours. For the Liberation of selenium, according to
the sample type used. Then take sodium hydroxide which brought together the element for
Determine the selenium in sample.

2.10. Preparation of the calibration graph

An aliquot of the stock solution containing 1.0-200 µM of selenium (IV) was transferred into
a 25 mL volumetric flask. Hydrochloric acid (10mL of 5M) and DAUH (2.5 mL of 10⁻³ mM) were
added. The solution was diluted up to the mark with distilled water and mixed well. The absorbance
of the solution was measured after about 30 minutes at 343 nm against a reagent blank. The amount
of selenium in the sample solution was then deduced from the calibration graph.

3. RESULTS AND DISCUSSION

Selenium reacts with 5,6-diaminouracil hydrochloride (DAUH) in hydrochloric acid solution
1:1 and gives yellow colored complex. The complex has a maximum absorbance at 343 nm. The
optimum reaction conditions for the quantitative determination of the metal-ligand complex was
established through a number of preliminary studies, such as the effect of acidic medium, reagent
concentration, interference of foreign ions, in order to develop a rapid, selective and sensitive
spectrophotometric method for the determination of selenium (IV) at microgram levels.

3.1. Absorption spectra of the reagent and Se (IV)-DAUH complex

Absorption spectra of Se (IV)-DAUH complex and reagent show maximum absorbance at 343 nm and 260 nm, respectively (Figure 3). The reagent showed minimum absorbance at the
wavelength of maximum absorbance of the complex. Hence, all the spectral measurements of the
complex were therefore carried out at 360 nm.

![Figure 3](image-url)

**Figure 3.** (1) Absorption spectra of reagent
(2) Absorption spectra of Se-DAUH complex
3.2. Optimization of reaction conditions

The spectrophotometric properties of the colored product as well as the different experimental parameters affecting the color development and its stability were studied and optimized by changing each variable in turn, while keeping all others constants. In all experiments.

3.3. Effect of hydrochloric acid, concentration.

From hydrochloric acid, phosphoric acid and sulphuric acid, the first gave the highest absorbance and stability. The study of the effect of acidity on the color intensity of the reaction mixture showed that the constant and maximum color is obtained in the range 1.6-2.8M, the. At concentrations higher than 2.8 M and lower than 1.6 M the absorbance decreases significantly. The working HCl concentration selected for the formation and fixation of the complex at 2M. The optimum HCl concentration is the same for all sample volumes studied. (Figure 4). The maximum intensity of the yellow color was achieved in the range of 5.5-10 ml concentrated hydrochloric acid. Therefore, 10 ml of hydrochloric acid 2M was used for the better results.

3.4. Effect of reagent concentration

All analytical studies were therefore, carried out at \(C_{\text{HCl}}=2.0\text{M}\). Different volume of molar excess of DAUH was added to fixed Se (IV) concentration and the absorbance's were measured adopting the standard procedure. It was observed that 100 fold molar excess of reagent with respect to metal ion is necessary to get maximum absorbance. Hence, a 100 fold molar excess of reagent was used for further experimental studies. The absorbance of the solution was measured at different time intervals to ascertain the time stability of the color complex.

![Absorption spectra in varying HCl molarity.](image1)

**Figure 4.** A - Absorption spectra in varying HCl molarity. B - the HCl molarity effect on the absorbance of the formed complex Se(IV)-DAUH.

3.5. Time, Temperature and Color Stability

Under the optimized conditions, although the color developed instantaneously, 30 min were allowed to obtain the maximum and constant absorbance in the method. The yellow colored product was stable for one week. The absorbance varied by not more than 1% over a period of two days for the method and color development was independent of temperature in the range of 25 °C – 45 °C

3.6. Analytical method validation

3.6.1. Calibration, graph reproducibility and detection limit

Using the optimized composition and conditions described above . The effect of the Se concentration was studied over 1.00-200.0 µM for convenience of the measurement. The calibration curves gave an excellent linear for (2.5-189.9µM), (0.2-15µg/mL), as shown in (Fig 5) at 343nm,
The molar absorption coefficient and the Shandell's sensitivity were found to be $0.5 \times 10^4$ L mol$^{-1}$ cm$^{-1}$ and 0.031 µg cm$^{-2}$ of Se respectively. The selected analytical parameters obtained are summarized in Table 1.

### 3.6.2. Precision and accuracy

The precision and accuracy of the method was studied by analysing solutions containing known amounts of Se (IV) within the Beer’s law limit. Percentage relative standard deviation (RSD %) as precision and percentage recovery as accuracy of the suggested method were calculated and showed in Table 2. The values of relative standard deviations for different concentrations of Se determined from the calibration curves. These results of accuracy and precision show that the proposed method have good repeatability and reproducibility. The lower values of relative standard deviation ( % ) and percentages of error indicated the high accuracy of the method.

![Absorption spectra in varying Se concentration](image)

**Figure 5.** - Absorption spectra in varying Se concentration

The linear range between absorbance formed complex and Se concentration

<table>
<thead>
<tr>
<th>Parameters</th>
<th>Values</th>
</tr>
</thead>
<tbody>
<tr>
<td>Wavelength / $\lambda_{\text{max}}$ (nm), complex</td>
<td>343</td>
</tr>
<tr>
<td>Wavelength / $\lambda_{\text{max}}$ (nm), reagent</td>
<td>260</td>
</tr>
<tr>
<td>Solvent</td>
<td>water</td>
</tr>
<tr>
<td>Time / min</td>
<td>30</td>
</tr>
<tr>
<td>Temperature /$^\circ$C</td>
<td>25±40$^\circ$C</td>
</tr>
<tr>
<td>Mole of reagent required mole of metal ion for full color developed</td>
<td>100 Fold</td>
</tr>
<tr>
<td>Composition of complex as obtained in Job’s and molar ratio methods (M:L)</td>
<td>1 : 1</td>
</tr>
<tr>
<td>Molar absorption Coefficient/ L mol$^{-1}$cm$^{-1}$</td>
<td>$0.5\times10^4$</td>
</tr>
<tr>
<td>Linear range/µg mL$^{-1}$</td>
<td>0.2-15</td>
</tr>
<tr>
<td>Detection limit/µg mL$^{-1}$</td>
<td>0.05</td>
</tr>
<tr>
<td>Sandell’s Sensitivity/µgcm$^{-2}$</td>
<td>0.031</td>
</tr>
<tr>
<td>Relative Standard Deviation</td>
<td>0.19-1.60</td>
</tr>
<tr>
<td>Regression Co-efficient</td>
<td>0.999</td>
</tr>
<tr>
<td>Slope</td>
<td>0.0099</td>
</tr>
</tbody>
</table>
Table 2. Accuracy and precision for the determination of Se in pure solution

<table>
<thead>
<tr>
<th>Taken</th>
<th>Found*</th>
<th>SD</th>
<th>RSD %</th>
<th>Confidence limit</th>
<th>ASE, µM</th>
<th>Recovery %</th>
</tr>
</thead>
<tbody>
<tr>
<td>10</td>
<td>9.86</td>
<td>0.16</td>
<td>1.60</td>
<td>9.86±0.20</td>
<td>0.07</td>
<td>98.57</td>
</tr>
<tr>
<td>30</td>
<td>28.71</td>
<td>0.22</td>
<td>0.75</td>
<td>28.71±0.27</td>
<td>0.10</td>
<td>95.69</td>
</tr>
<tr>
<td>50</td>
<td>47.63</td>
<td>0.20</td>
<td>0.43</td>
<td>47.63±0.25</td>
<td>0.09</td>
<td>95.27</td>
</tr>
<tr>
<td>70</td>
<td>69.90</td>
<td>0.19</td>
<td>0.28</td>
<td>69.90±0.24</td>
<td>0.09</td>
<td>99.86</td>
</tr>
<tr>
<td>100</td>
<td>104.21</td>
<td>0.19</td>
<td>0.19</td>
<td>104.21±0.24</td>
<td>0.09</td>
<td>104.21</td>
</tr>
</tbody>
</table>

aFive independent analyses.

3.6.3. Effect of diverse ions

The extent of interference by diverse ions was determined by measuring the absorbance of solutions containing 5 µg/mL of Se (IV) and various amounts of diverse ions. The criterion for an interference was an absorbance value varying by more than ±2% from the expected value of selenium (IV) alone. The results presented in Table 3 show that a large excess of cations and anions which are usually associated in the determination of Se (IV), do not interfere.

Table 3. Tolerance limits of diverse ions in the determination of 5 µg/mL of selenium (IV).

<table>
<thead>
<tr>
<th>Ion Added</th>
<th>Tolerance limit µg/mL</th>
<th>Ion Added</th>
<th>Tolerance limit µg/mL</th>
</tr>
</thead>
<tbody>
<tr>
<td>Fe(II)</td>
<td>125</td>
<td>Cd (II)</td>
<td>65</td>
</tr>
<tr>
<td>Fe(III)</td>
<td>75</td>
<td>Ag (I)</td>
<td>225</td>
</tr>
<tr>
<td>Hg(II)</td>
<td>160</td>
<td>V (V)</td>
<td>310</td>
</tr>
<tr>
<td>Ni(II)</td>
<td>300</td>
<td>Tartrate</td>
<td>485</td>
</tr>
<tr>
<td>Zn(II)</td>
<td>420</td>
<td>Iodate</td>
<td>302</td>
</tr>
<tr>
<td>Mn(II)</td>
<td>285</td>
<td>Urea</td>
<td>288</td>
</tr>
<tr>
<td>Cr(III)</td>
<td>300</td>
<td>Citrate</td>
<td>386</td>
</tr>
<tr>
<td>Ca(II)</td>
<td>400</td>
<td>Bicarbonate</td>
<td>423</td>
</tr>
<tr>
<td>Ba(II)</td>
<td>380</td>
<td>Sulphate</td>
<td>368</td>
</tr>
<tr>
<td>Cu(II)</td>
<td>600</td>
<td>Oxalate</td>
<td>257</td>
</tr>
<tr>
<td>Ni(II)</td>
<td>740</td>
<td>Nitrate</td>
<td>320</td>
</tr>
<tr>
<td>Pb (II)</td>
<td>8.3</td>
<td>Acetate</td>
<td>189</td>
</tr>
</tbody>
</table>

3.6.4. Composition and stability constant of the complex

Job’s method of continuous variation and molar-ration methods were applied to ascertain the stoichiometric composition of the complex. It was found that DAUH forms 1:1 complex with Se (IV) as shown in the (Fig.6).

Figure 6. Continuous variation and molar-ration methods. Se (IV)-DAUH, Se (IV) and DAUH,1 ×10⁻³ M
3.6.5. Application

The proposed spectrophotometric method is applied for the determination of Se (IV) in various samples. A known aliquot of the above sample solutions were taken and the selenium content was determined as described is given in the general procedure, and the results of the suggested method were coincidental with the analysis data of the same samples with the Hydride Generation-Atomic Absorption technique as a comparative method. Table (4).

Table 4: Determination Se(IV) in various samples

<table>
<thead>
<tr>
<th>Samples</th>
<th>Results of HG-AAS</th>
<th>Results of proposed method</th>
</tr>
</thead>
<tbody>
<tr>
<td></td>
<td>$\bar{X}$ ± $\Delta X$, ppm</td>
<td>RSD %</td>
</tr>
<tr>
<td>Water</td>
<td>0.534±0.032</td>
<td>1.85</td>
</tr>
<tr>
<td>Yellow Corn</td>
<td>0.691±0.019</td>
<td>2.20</td>
</tr>
<tr>
<td>Peas</td>
<td>0.730±0.011</td>
<td>1.20</td>
</tr>
<tr>
<td>Chick-pea</td>
<td>0.782±0.009</td>
<td>0.90</td>
</tr>
<tr>
<td>Wheat</td>
<td>0.950±0.005</td>
<td>0.40</td>
</tr>
<tr>
<td>Almond</td>
<td>1.190±0.022</td>
<td>1.50</td>
</tr>
<tr>
<td>Concentrated forage of chickens (1)</td>
<td>0.890±0.002</td>
<td>0.20</td>
</tr>
<tr>
<td>Concentrated forage of chickens (2)</td>
<td>0.991±0.012</td>
<td>1.00</td>
</tr>
<tr>
<td>Over-concentrated forage of chickens</td>
<td>16.000±0.654</td>
<td>3.30</td>
</tr>
<tr>
<td>Stannous metal (Egg hen)</td>
<td>19.850±0.246</td>
<td>1.00</td>
</tr>
<tr>
<td>Syravite E (Liquid solution)</td>
<td>0.670±0.012</td>
<td>1.40</td>
</tr>
<tr>
<td>Syravite E + Selenium powder</td>
<td>0.811±0.007</td>
<td>0.70</td>
</tr>
</tbody>
</table>

4. CONCLUSION

The author has introduced a new sensitive reagent DAUH for the direct spectrophotometric determination of trace amounts of Se (IV). The proposed spectrophotometric method is simple, highly sensitive and selective for the determination of Se (IV) in water and soil samples when compared with other spectrophotometric methods. The proposed method is simple, rapid and common metal ions such as Fe$^{3+}$, Pb$^{2+}$, Co$^{2+}$, Ni$^{2+}$, Zn$^{2+}$, Mn$^{2+}$, Cr$^{3+}$ do not interfere. It also offers advantages like reliability and reproducibility in addition to its simplicity instant color development and less interference effect. The results obtained through UV-Visible spectrophotometer have been compared with those obtained through the HG-AAS. The method has been successfully applied for the determination of selenium in various samples.

References