

Why hot water freezes faster than cold?

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ABSTRACT

The modern science mainly treats the biochemical basis of sequencing in bio-macromolecules and processes in chemistry and biochemistry. In these informational cybernetic principles we seek to answer the following question: Why hot water freezes faster than cold? These principles determine the freezing point of water. Determine that the hot water freezes faster than cold. What we did is the following: We translated the physical and chemical parameters from the language of water into the digital language of programmatic, cybernetic and information principles. This we did by using the adequate mathematical algorithms. By using chemical-information procedures, we calculated the numerical value for the information content of molecule of water. What we got this way is the digital picture of this molecule. These digital pictures reveal to us a whole new dimension of water. They reveal to us that the chemical process of this molecule is strictly conditioned and determined by programmatic, cybernetic and information principles.

Keywords: Water; Hot Water; Freezing hot water; Freezing cold water

1. INTRODUCTION

In these informational cybernetic principles we seek to answer the following question: Why hot water freezes faster than cold?

These principles determine the freezing point of water. Determine that the hot water freezes faster than cold. What we did is the following: We translated the physical and chemical parameters from the language of water into the digital language of programmatic, cybernetic and information principles. This we did by using the adequate mathematical algorithms.

By using chemical-information procedures, we calculated the numerical value for the information content of molecule of water.

What we got this way is the digital picture of this molecule. These digital pictures reveal to us a whole new dimension of water. They reveal to us that the chemical process of this molecule is strictly conditioned and determined by programmatic, cybernetic and information principles.

2. RESULTS

Digital image where chemistry is represented is in the form of numbers. That image can be created with the help of the new scientific methods. At the first stage of our research we replaced the chemical elements from the chemical formulae with atomic numbers, numbers of atoms, atomic weight and other numerical values in those formulas.

By this way we got digital image of chemistry. Then we mathematically analyzed those digital images. After we making such analysis, we discovered existence of digital codes in those images, which interconnect all chemical elements and other sequences in chemistry. Given below is a brief introduction about the way we discovered those chemistry digital codes and how those codes interconnect all chemical elements in this science.

2. 1. Dissociation constant of water

As a frequently used special case, the dissociation constant of water is often expressed as K_w : The concentration of water is not included in the definition of K_w , for reasons described in the article equilibrium constant.

The value of K_w varies with temperature, as shown in the table below. This variation must be taken into account when making precise measurements of quantities such as pH.

Why hot water freezes faster than cold?

Water temperature	$K_w / 10^{-14}$	pK_w
0 °C	0.1	14.92
10 °C	0.3	14.52
18 °C	0.7	14.16
25 °C	1.2	13.92
30 °C	1.8	13.75
50 °C	8.0	13,10
60 °C	12.6	12.90
70 °C	21.2	12.67
80 °C	35	12.46
90 °C	53	12.28
100 °C	73	12.14

2. 2. Digital language of dissociation constant of water

The matrix mechanism of water temperature have been analyzed by the application of cybernetic methods, information theory and system theory, respectively. First, we decode the dissociation constant of water expressed as pK_w :

Example 1.

Water temperature	$K_w / 10^{-14}$	pK_w
0 °C	0.1	14.92
100 °C	73	12.14
Sum		27,06

Water temperature	$K_w / 10^{-14}$	pK_w
18 °C	0.7	14.16
60 °C	12.6	12.90
Sum		27,06

$$[(0 + 100) - (60 + 18)] = 22; (27,06 \times 100) = (22 + 22 + 22 \dots + 22);$$

Determinante 2 x 2

$$\text{DET}(0.1, 14.92, 12.14, 73) = \mathbf{-168};$$

$$\text{DET}(0.7, 14.16, 12.6, 12.90) = \mathbf{-168};$$

In this example there a mathematical balance between constants $K_w / 10^{-14}$ and pK_w . This balance of the determine program, cyber laws and information.

Example 2.

Water temperature	$K_w / 10^{-14}$	pK_w
10 °C	0.3	14.52
18 °C	0.7	14.16
60 °C	12.6	12.90
90 °C	53	12.28
Sum		53,86

Water temperature	$K_w / 10^{-14}$	pK_w
0 °C	0.1	14.92
10 °C	0.3	14.52
90 °C	53	12.28
100 °C	73	12.14
Sum		53,86

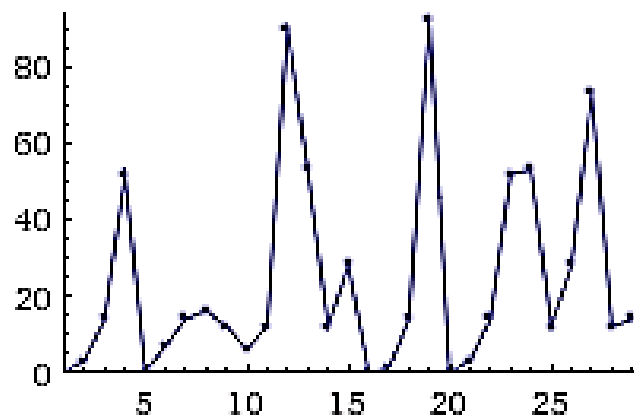
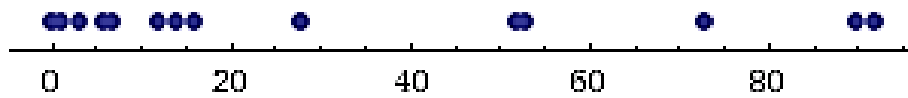
Determinate 4 x 4

0,3	14,52	0,7	14,16
12,6	12,90	53	12,28
0,1	14,92	0,3	14,52
53	12,28	73	12,14

↓

0**Input:**

{0, 3, 14, 52, 0, 7, 14, 16, 12, 6, 12, 90, 53,
12, 28, 0, 1, 14, 92, 0, 3, 14, 52, 53, 12, 28, 73, 12, 14}

Plot:**Number line:****Length of data:**

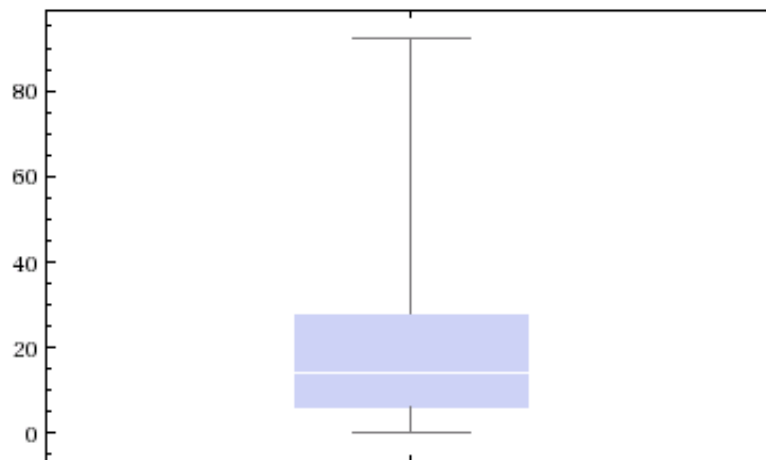
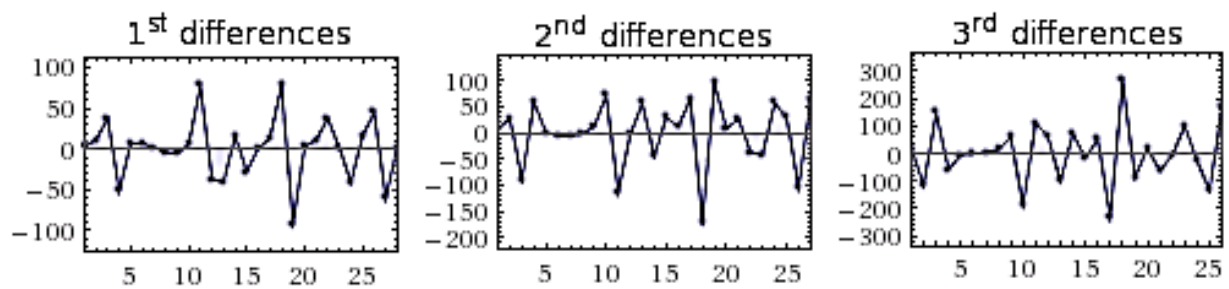
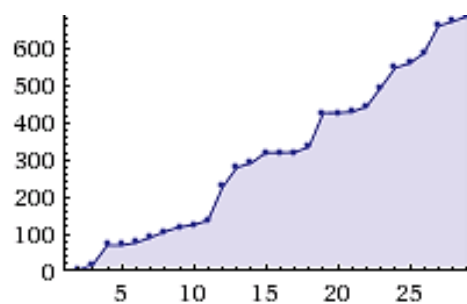
29 items

Total:

$$0 + 3 + \dots + 12 + 14 = 687$$

Statistics:

mean	23.69
median	14
sample standard deviation	26.91

Box-and-whisker chart:**Differences:****Cumulative sums:**

In this example there is a balance of size $K_w / 10^{-14}$ and pK_w .

Example 3.

Water temperature	$K_w / 10^{-14}$	pK_w
10 °C	0.3	14.52
18 °C	0.7	14.16
60 °C	12.6	12.90
80 °C	35	12.46
Sum		54,04

Water temperature	$K_w / 10^{-14}$	pK_w
0 °C	0.1	14.92
10 °C	0.3	14.52
80 °C	35	12.46
100 °C	73	12.14
Sum		54,04

Example 4.

Water temperature	$K_w / 10^{-14}$	pK_w
18 °C	0.7	14.16
25 °C	1.2	13.92
30 °C	1.8	13.75
50 °C	8.0	13,10
60 °C	12.6	12.90
Sum		67,83

Water temperature	$K_w / 10^{-14}$	pK_w
0 °C	0.1	14.92
25 °C	1.2	13.92
30 °C	1.8	13.75
50 °C	8.0	13,10
100 °C	73	12.14
Sum		67,83

Example 5.

Water temperature	$K_w / 10^{-14}$	pK_w
18 °C	0.7	14.16
25 °C	1.2	13.92
30 °C	1.8	13.75
50 °C	8.0	13,10
60 °C	12.6	12.90
70 °C	21.2	12.67
Sum		80,50

Water temperature	$K_w / 10^{-14}$	pK_w
0 °C	0.1	14.92
25 °C	1.2	13.92
30 °C	1.8	13.75
50 °C	8.0	13,10
70 °C	21.2	12.67
100 °C	73	12.14
Sum		80,50

Water temperature	$K_w / 10^{-14}$	pK_w
0 °C	0.1	14.92
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25 °C	1.2	13.92
60 °C	12.6	12.90
80 °C	35	12.46
100 °C	73	12.14
Sum		80,50

Example 6.

Water temperature	$K_w / 10^{-14}$	pK_w
18 °C	0.7	14.16
25 °C	1.2	13.92
30 °C	1.8	13.75
50 °C	8.0	13,10
60 °C	12.6	12.90
70 °C	21.2	12.67
80 °C	35	12.46
Sum		92,96

Water temperature	$K_w / 10^{-14}$	pK_w
0 °C	0.1	14.92
25 °C	1.2	13.92
30 °C	1.8	13.75
50 °C	8.0	13,10
70 °C	21.2	12.67
80 °C	35	12.46
100 °C	73	12.14
Sum		92,96

etc.

From the above examples we see that the dissociation constant of water determine the programmatic cyber laws. Here are some examples:

Example 1.

Water temperature	$K_w / 10^{-14}$
80 °C	35
Sum	35

Water temperature	$K_w / 10^{-14}$
25 °C	1.2
60 °C	12.6
70 °C	21.2
Sum	35

Example 2.

Water temperature	$K_w / 10^{-14}$	pK_w
0 °C	0.1	14.92
100 °C	73	12.14

↓

Determinate

14,92	0
100	12.14

$$\text{DET}(14,92, 0, 100, 12,14) = \mathbf{1400};$$

Example 3.

Water temperature	$K_w / 10^{-14}$	pK_w
10 °C	0.3	14.52
90 °C	53	12.28

↓

Determinate

14,52	10
90	12.28

$$\text{DET}(14,52, 10, 90, 12,28) = \mathbf{1140};$$

etc.

In this example there is also a mathematical balance between constants pK_w and water temperature. This balance is one of important quantitative characteristics of all processes in chemistry.

2. 3. Matrix in the water molecule

We shall now give some mathematical evidences that will prove that in the chemistry of water there really is programmatic and cybernetic algorithm in which it is „recorded“, in the language of mathematics, how the molecule will be built and what will be the quantitative characteristics of the given chemical information.

Water temperature	K_w
0 °C	0.1
10 °C	0.3
18 °C	0.7
25 °C	1.2
30 °C	1.8
50 °C	8.0
60 °C	12.6
70 °C	21.2
80 °C	35
90 °C	53
100 °C	73

This table contains an overview of constants K_w . The values show some of the quantitative characteristics of the molecule of water. Actually, they show that there is an exact mathematical balance between these constants. Examples:

	Cumulative										
K_w	0,1	0,3	0,7	1,2	1,8	8,0	12,6	21,2	35,0	53,0	73,0
AP-a	0,1	0,4	1,1	2,3	4,1	12,1	24,7	45,9	80,9	133,9	206,9
AP-b	206,9	206,8	206,5	205,8	204,6	202,8	194,8	182,2	161,0	126,0	73,0

$(0,1+0,3) = 0,4$; $(0,4+0,7) = 1,1$; $(1,1+1,2) = 2,3$;
etc.

	Matrix AP-a										
206,9	206,8	206,5	205,8	204,6	202,8	194,8	182,2	161,0	126,0	73,0	
133,9	0	133,8	133,5	132,8	131,6	129,8	121,8	109,2	88,0	53,0	
80,9	0	0	80,8	80,5	79,8	78,6	76,8	68,8	56,2	35,0	
45,9	0	0	0	45,8	45,5	44,8	43,6	41,8	33,8	21,2	
24,7	0	0	0	0	24,6	24,3	23,6	22,4	20,6	12,6	
12,1	0	0	0	0	0	12,0	11,7	11,0	9,8	8,0	
4,1	0	0	0	0	0	0	4,0	3,7	3,0	1,8	
2,3	0	0	0	0	0	0	0	2,2	1,9	1,2	
1,1	0	0	0	0	0	0	0	0	1,0	0,7	
0,4	0	0	0	0	0	0	0	0	0	0,3	
0,1	0	0	0	0	0	0	0	0	0	0	
206,8	340,3	420,1	463,7	484,3	484,3	463,7	420,1	340,3	206,8		

$(206,9-0,1) = 206,8$; $(206,9-0,4) = 206,5$; ; $(206,9-1,1) = 205,8$;
etc.

This is one of the digital image water molecules. This picture reveals the secret of digital connection between water temperature and constant K_w . Also, he reveals a balance of mathematical constants. This balance has emerged as a result of chemical processes. It is a result of software, information and cyber laws. Cumulative from the Matrix AP-a are in correlation with each other.

Matrix inversion AP-a.

206,9	0	0	0	0	0	0	0	0	0	0	0
133,9	0	-73,0	0	0	0	0	0	0	0	0	0
80,9	0	-53,0	-126,0	0	0	0	0	0	0	0	0
45,9	0	-35,0	-88,0	-161,0	0	0	0	0	0	0	0
24,7	0	-21,2	-56,2	-109,2	-182,2	0	0	0	0	0	0
12,1	0	-12,6	-33,8	-68,8	-121,8	-194,8	0	0	0	0	0
4,1	0	-8,0	-20,6	-41,8	-76,8	-129,8	-202,8	0	0	0	0
2,3	0	-1,8	-9,8	-22,4	-43,6	-78,6	-131,6	-204,6	0	0	0
1,1	0	-1,2	-3,0	-11,0	-23,6	-44,8	-79,8	-132,8	-205,8	0	0
0,4	0	-0,7	-1,9	-3,7	-11,7	-24,3	-45,5	-80,5	-133,5	-206,5	0
0,1	0	-0,3	-1,0	-2,2	-4,0	-12,0	-24,6	-45,8	-80,8	-133,8	-206,8
Sum		-206,8	-340,3	-420,1	-463,7	-484,3	-484,3	-463,7	-420,1	-340,3	-206,8

Inversion matrix AP is analogous to the matrix of the matrix AP-a. Her task is discreet in its memory keeps information from this matrix.

Matrix AP-b

206,9	133,9	80,9	45,9	24,7	12,1	4,1	2,3	1,1	0,4	0,1
206,8	0	133,8	80,8	45,8	24,6	12,0	4,0	2,2	1,0	0,3
206,5	0	0	133,5	80,5	45,5	24,3	11,7	3,7	1,9	0,7
205,8	0	0	0	132,8	79,8	44,8	23,6	11,0	3,0	1,2
204,6	0	0	0	0	131,6	78,6	43,6	22,4	9,8	1,8
202,8	0	0	0	0	0	129,8	76,8	41,8	20,6	8,0
194,8	0	0	0	0	0	0	121,8	68,8	33,8	12,6
182,2	0	0	0	0	0	0	0	109,2	56,2	21,2
161,0	0	0	0	0	0	0	0	0	88,0	35,0
126,0	0	0	0	0	0	0	0	0	0	53,0
73,0	0	0	0	0	0	0	0	0	0	0
	133,9	214,7	260,2	283,8	293,6	293,6	283,8	260,2	214,7	133,9

In this matrix are stored information about the constants K_w . The matrices AP- a and AP-b was determined by the constant K_w , must have water temperatures in this example.

Matrix inversion AP-b.

206,9	0	0	0	0	0	0	0	0	0	0	0
206,8	0	-0,1	0	0	0	0	0	0	0	0	0
206,5	0	-0,3	-0,4	0	0	0	0	0	0	0	0
205,8	0	-0,7	-1,0	-1,1	0	0	0	0	0	0	0
204,6	0	-1,2	-1,9	-2,2	-2,3	0	0	0	0	0	0
202,8	0	-1,8	-3,0	-3,7	-4,0	-4,1	0	0	0	0	0
194,8	0	-8,0	-9,8	-11,0	-11,7	-12,0	-12,1	0	0	0	0
182,2	0	-12,6	-20,6	-22,4	-23,6	-24,3	-24,6	-24,7	0	0	0
161,0	0	-21,2	-33,8	-41,8	-43,6	-44,8	-45,5	-45,8	-45,9	0	0
126,0	0	-35,0	-56,2	-68,8	-76,8	-78,6	-79,8	-80,5	-80,8	-80,9	0
73,0	0	-53,0	-88,0	-109,2	-121,8	-129,8	-131,6	-132,8	-133,5	-133,8	-133,9
						-					
						-293,6	293,6				
								-283,8	-260,2	-214,7	-133,9

Differences in the mean number totals matrices AP and AP-B associated with the 1907. This code connects the molecules in nature.

$$[(484,3 - 293,6) \times 100] = 1907;$$

Constant pK_w.

Water temperature	pK _w
0 °C	14.92
10 °C	14.52
18 °C	14.16
25 °C	13.92
30 °C	13.75
50 °C	13,10
60 °C	12.90
70 °C	12.67
80 °C	12.46
90 °C	12.28
100 °C	12.14

This table contains an overview of constant pK_w. The values show some of the quantitative characteristics of the molecule of water. Actually, they show that there is an exact mathematical balance between these constants. Examples:

Cumulative

	12,14	12,28	12,46	12,67	12,90	13,10	13,75	13,92	14,16	14,52	14,92
APa	12,14	24,42	36,88	49,55	62,45	75,55	89,30	103,22	117,38	131,90	146,82
APb	146,82	134,68	122,40	109,94	97,27	84,37	71,27	57,52	43,60	29,44	14,92

Matrix AP-a

146,82	134,68	122,40	109,94	97,27	84,37	71,27	57,52	43,60	29,44	14,92
131,90	0	119,76	107,48	95,02	82,35	69,45	56,35	42,60	28,68	14,52
117,38	0	0	105,24	92,96	80,50	67,83	54,93	41,83	28,08	14,16
103,22	0	0	0	91,08	78,80	66,34	53,67	40,77	27,67	13,92
89,30	0	0	0	0	77,16	64,88	52,42	39,75	26,85	13,75
75,55	0	0	0	0	0	63,41	51,13	38,67	26,00	13,10
62,45	0	0	0	0	0	0	50,31	38,03	25,57	12,90
49,55	0	0	0	0	0	0	0	37,41	25,13	12,67
36,88	0	0	0	0	0	0	0	0	24,74	12,46
24,42	0	0	0	0	0	0	0	0	0	12,28
12,14	0	0	0	0	0	0	0	0	0	0
Sum	134,68	242,16	322,66	376,33	403,18	403,18	376,33	322,66	242,16	134,68

Matrix inversion AP-a

146,82	0	0	0	0	0	0	0	0	0	0	0
131,90	0	-14,92	0	0	0	0	0	0	0	0	0
117,38	0	-14,52	-29,44	0	0	0	0	0	0	0	0
103,22	0	-14,16	-28,68	-43,60	0	0	0	0	0	0	0
89,30	0	-13,92	-28,08	-42,60	-57,52	0	0	0	0	0	0
75,55	0	-13,75	-27,67	-41,83	-56,35	-71,27	0	0	0	0	0
62,45	0	-13,10	-26,85	-40,77	-54,93	-69,45	-84,37	0	0	0	0
49,55	0	-12,90	-26,00	-39,75	-53,67	-67,83	-82,35	-97,27	0	0	0
36,88	0	-12,67	-25,57	-38,67	-52,42	-66,34	-80,50	-95,02	-109,94	0	0
24,42	0	-12,46	-25,13	-38,03	-51,13	-64,88	-78,80	-92,96	-107,48	-122,40	0
12,14	0	-12,28	-24,74	-37,41	-50,31	-63,41	-77,16	-91,08	-105,24	-119,76	-134,68
Sum	-134,68	-242,16	-322,66	-376,33	-403,18	-403,18	-376,33	-322,66	-242,16	-134,68	

Matrix AP-b

146,82	131,90	117,38	103,22	89,30	75,55	62,45	49,55	36,88	24,42	12,14
134,68	0	119,76	105,24	91,08	77,16	63,41	50,31	37,41	24,74	12,28
122,40	0	0	107,48	92,96	78,80	64,88	51,13	38,03	25,13	12,46
109,94	0	0	0	95,02	80,50	66,34	52,42	38,67	25,57	12,67
97,27	0	0	0	0	82,35	67,83	53,67	39,75	26,00	12,90
84,37	0	0	0	0	0	69,45	54,93	40,77	26,85	13,10
71,27	0	0	0	0	0	0	56,35	41,83	27,67	13,75
57,52	0	0	0	0	0	0	0	42,60	28,08	13,92
43,60	0	0	0	0	0	0	0	0	28,68	14,16
29,44	0	0	0	0	0	0	0	0	0	14,52
14,92	0	0	0	0	0	0	0	0	0	0
Sum	131,90	237,14	315,94	368,36	394,36	394,36	368,36	315,94	237,14	131,90

Matrix inversion AP-b

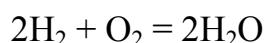
146,82	0	0	0	0	0	0	0	0	0	0	0	0
134,68	0	-12,14	0	0	0	0	0	0	0	0	0	0
122,40	0	-12,28	-24,42	0	0	0	0	0	0	0	0	0
109,94	0	-12,46	-24,74	-36,88	0	0	0	0	0	0	0	0
97,27	0	-12,67	-25,13	-37,41	-49,55	0	0	0	0	0	0	0
84,37	0	-12,90	-25,57	-38,03	-50,31	-62,45	0	0	0	0	0	0
71,27	0	-13,10	-26,00	-38,67	-51,13	-63,41	-75,55	0	0	0	0	0
57,52	0	-13,75	-26,85	-39,75	-52,42	-64,88	-77,16	-89,30	0	0	0	0
43,60	0	-13,92	-27,67	-40,77	-53,67	-66,34	-78,80	-91,08	-103,22	0	0	0
29,44	0	-14,16	-28,08	-41,83	-54,93	-67,83	-80,50	-92,96	-105,24	-117,38	0	0
14,92	0	-14,52	-28,68	-42,60	-56,35	-69,45	-82,35	-95,02	-107,48	-119,76	-131,90	0
Sum		-131,90	-237,14	-315,94	-368,36	-394,36	-394,36	-368,36	-315,94	-237,14	-131,90	0

Now we have the answer to the question: Why hot water freezes faster than cold? The answer is as follows: Cybernetic information principles in your digital program to determine the hot water freezes faster than cold water. It is written in the nature of the algorithm. Nature has decided that this is so. Is it really so testify digital image molecule of water. These pictures have countless. We mentioned this in just a few of these images. We hope that researchers in the field of chemistry will soon decode and other digital images.

2. 4. Chemical equilibrium

Under the law of mass action of reaction product is proportional to the concentration of substances. Under the law of mass action of reaction product is proportional to the concentration of substances. At temperatures of 60-10000 hydrogen and oxygen to combine with the explosion, and produced water, and water temperatures of 4000-50000 with an explosion breaks down into hydrogen and oxygen.

The reaction between hydrogen and oxygen is reversible.



In this way, chemical equilibrium is established. So, is done and the creation of water molecules and their decomposition. Numerical value of K determines the position of equilibrium at a given temperature. Therefore, the constant K preheated water has a numerical value, and the other with hot water, so the warmer water freezes faster than in the preheated. In this way the rapid freezing of water establishes the position of equilibrium hydrogen and oxygen at a given temperature. The equilibrium constant is calculated as follows:

$$K1 = [\text{H}_2\text{O}]^2$$

$$K2 = [\text{H}_2]^2 [\text{O}_2]$$

$$\text{mn}(K1 : K2) = K$$

The principle of moving equilibrium (Lešateljce, 1884): If the system that is in balance, made apoljni impacts, this balance shifts in the direction of influence and so long as the opposite effect, which is more and more increases in the system, not to equalize with the outside influence. Therefore, warm water freezes faster than preheated.

The equation shows that $2\text{H}_2 + \text{O}_2 = 2\text{H}_2\text{O} + 137 \text{ kcal}$ is incorporating hydrogen with oxygen followed by separation of heat and the decomposition of water by absorbing heat. If more heat is brought into the system, so much more beneficial effect on the decomposition of water, is the endothermic reaction.

3. CONCLUSION

Conversely, withdrawal of heat by cooling complicate the decomposition of water (exothermic reaction). If this system at a certain temperature is in equilibrium, and then warms up, the balance will be appropriately moved to the side creating a larger concentration of free hydrogen and oxygen.

So, when heating a balanced system, and this water is preheated, the balance shifts to the side of endothermic reactions, and on cooling the side exothermic. Thus, hot water freezes faster than preheated to establish a balance of hydrogen and oxygen at a given temperature. The equation shows that X is incorporating hydrogen with oxygen followed by separation of heat and the decomposition of water by absorbing heat.

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